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GASKELL Problem

1.1* The plot of $V = V(P, T)$ for a gas is shown in Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these

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(second derivatives).

Solution

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DH . Work is found
thethe first law as $w = q$
- DU ; thus $q=DH$;
 $w=D HPVL$; 4.

Isothermal Process

Because U is a function

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only of T for an ideal gas, $DU = DH = 0$ for an isothermal process.

These results also follow from the general results by using $DT = D(PV) = 0$ for an isothermal process.

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GASKELL Problem

2.8* One mole of a monatomic ideal gas at standard temperature and pressure (STP) undergoes the following three processes: 1. at constant Pressure the Temperature is

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thus $q = DH$; $w = D$

HPVL; 4. Isothermal

Process Because U is a

function only of T for

an. ideal gas, $DU = DH$

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The three key variables are pressure, P , temperature, T , and volume, V . It has been observed that when P and T are fixed that V always has a unique value. In other words, P and T are the independent variables and V is a function of P and T : $\text{Volume} = V@P, T$; Such an equation

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is called an equation of
state.

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Isothermal Process

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[134wwr85xw47] Work is found the the first law

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